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1. Elements with large electron affinities tend to form monoatomic anions  
2. No ionic compounds are found with positive ions having charges greater than the element group number  
3. Group 1A and 2A metals

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always have a positive charge equal to their group number in their ionic compounds.

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Substitution and  
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statement that uses  
chemical formulas to  
show the identities...  
The process by which  
the atoms of one or  
more substances are r...  
Substances formed  
during the reaction. A  
statement that uses  
chemical formulas to  
show the identities....

**chemistry chapter 9  
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Answers may include release or absorption of energy, change in color, change in odor, formation of a gas, or formation of a solid. 9. Compare and contrast a skeleton equation and a chemical equation. The skeleton equation includes the formulas of reactants and products. The chemical equation gives the relative



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amounts of reactants and products. 10.

## **Chemical Reactions Chemical Reactions**

Chapter 9 Chemical Quantities. 1. Although we define mass as the “amount of matter in a substance,” the units in which we measure mass are a human invention. Atoms and molecules react on an individual particle- by- particle basis, and we

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Answers  
have to count  
individual particles  
when doing chemical  
calculations.

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A polymer in which at  
least 100 amino acid  
monomers are linked  
through bonds  
between an amino  
group and a carboxyl

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group. photosynthesis.  
The process during  
which plants  
chemically combine  
carbon dioxide and  
water into  
carbohydrates.

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the O — O bond  
distance is relatively  
short. there are two  
unpaired electrons in  
the MO electron  
configuration of O 2 .

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the energy of the  $\pi$  2p MOs is higher than that of the  $\sigma$  2p MO. there are more electrons in the bonding orbitals than in the antibonding orbitals. the bond order in  $O_2$  can be shown to be equal to 2.

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textbook, read about the activity series for metals and halogens. II on Examine each of the following pairs of potential reactants. Use Figure 9.13 in your textbook to help you decide whether or not a reaction would occur. If a reaction occurs, write the balanced equation.

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Biological Chemistry.

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Chapter 9 Chemical Calculations and Chemical Formulas  
329 n order to explore and make use of the seemingly limitless changes that matter can undergo, chemists and chemistry students often need to answer questions that

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18.9 Occurrence,  
Preparation, and  
Compounds of Oxygen

18.10 Occurrence,  
Preparation, and  
Properties of Sulfur

18.11 Occurrence,  
Preparation, and  
Properties of Halogens

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Law Use the periodic  
table below to answer  
the questions in this

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Chapter Test. In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

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This video explains the concepts from your packet on Chapter 9 (Molecular Geometry and Bonding Theories), which can be found

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## **Chapter 9 Molecular Geometry and Bonding Theories**

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Reactions Chapter  
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reactions corresponds  
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...

After reading Lesson 9.4, answer the following questions.

Names and Formulas of Acids hydrogen1. Acids produce ions when dissolved in water.

anions2. When naming acids, you can consider them to be combinations of connected to as many ions as are necessary to create an electrically



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neutral compound. 3.

HBr What is the formula  
for hydrobromic acid?

4.

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